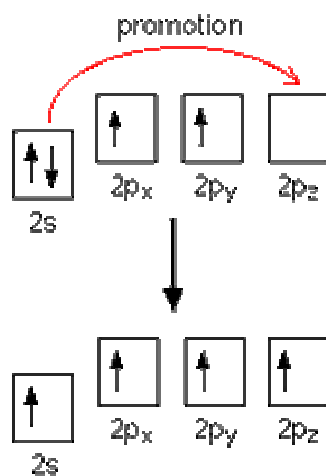


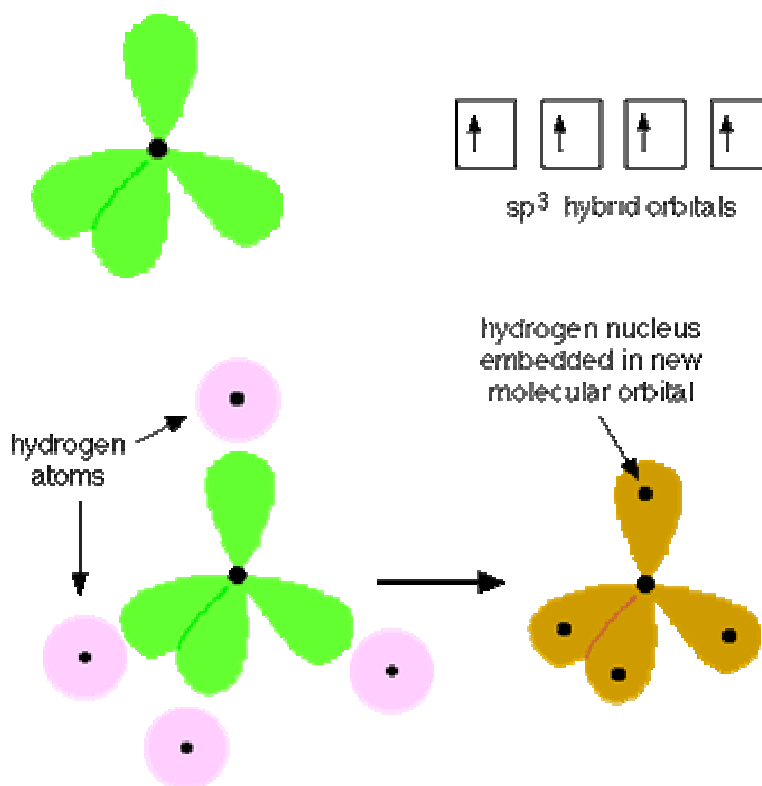
Metan, CH₄

sp³ hibridizacija, σ veze

Promocija → hibridizacija

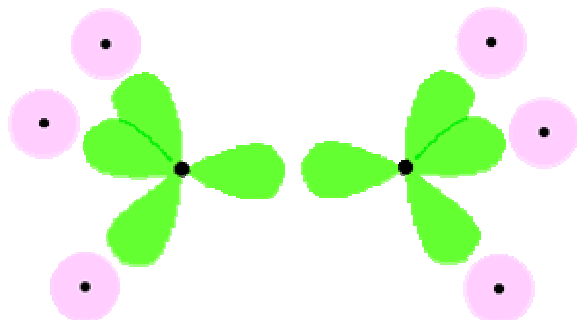


Hibridizacija: elektroni se reorganizuju u 4 identične hibridne orbitale, sp³ orbitale; 109,47°

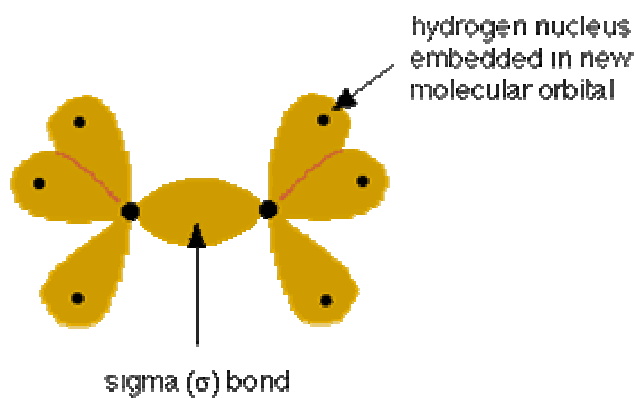


Etan, C₂H₆

sp³ hibridizacija, σ veze

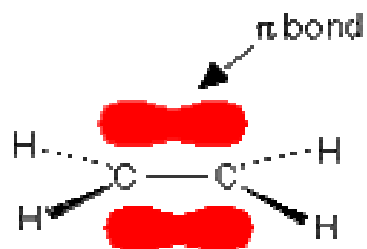
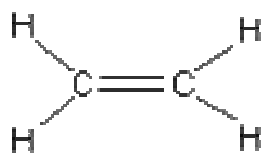
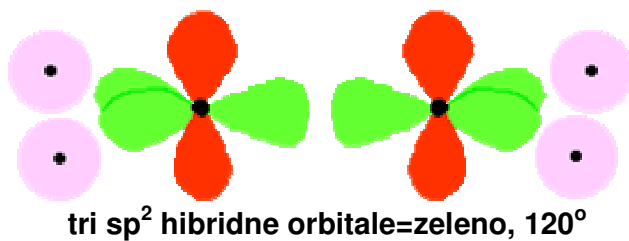


četiri sp³ orbitale= zeleno



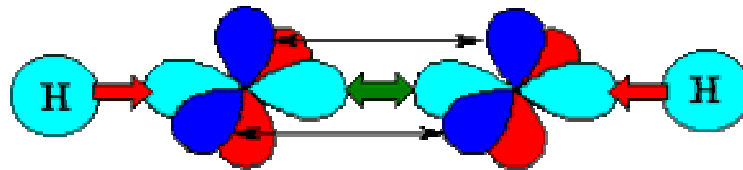
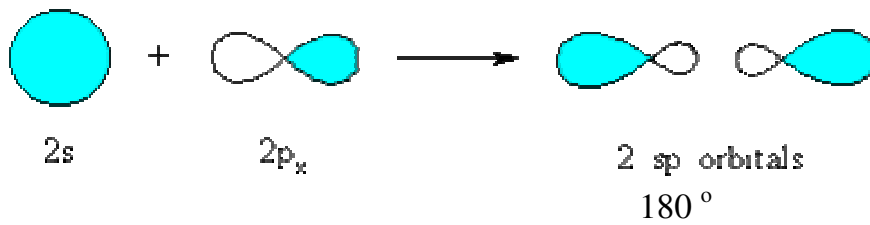
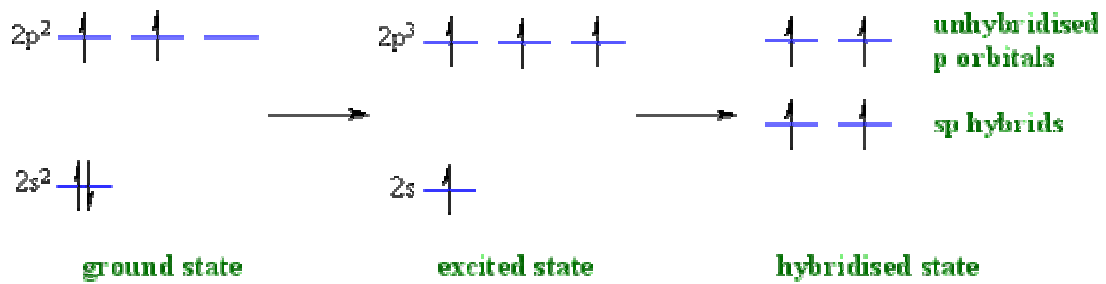
Eten, C₂H₄

sp² hibridizacija, σ i π veze, dvostruka veza



Etin, C₂H₂

sp hibridizacija, σ i π veze, trostruka veza



- o oba C su sp hibridizovana
- o 2 C-H σ veze su napravljene interakcijom C sp sa H1s orbitalama (videti **crvene** strelice)
- o 1 C-C σ veza je napravljena interakcijom C sp sa drugom C sp orbitalom (videti **zelenu** strelicu)

2 C-C π veze su izgrađene interakcijom 2 para C p orbitals (videti **crne** strelice).

Trostruka veza: jedna σ veza + dve π veze